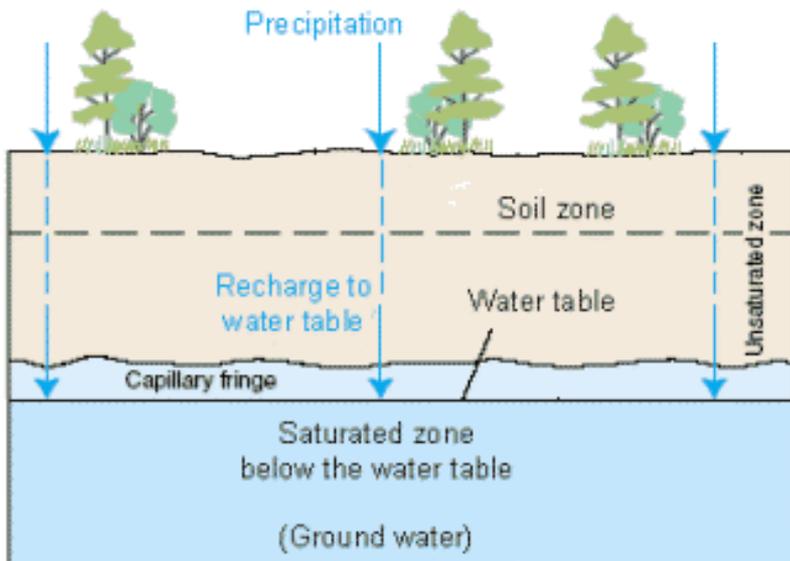


# **Usage of Water-Filled Trench in Improving Groundwater Quality**

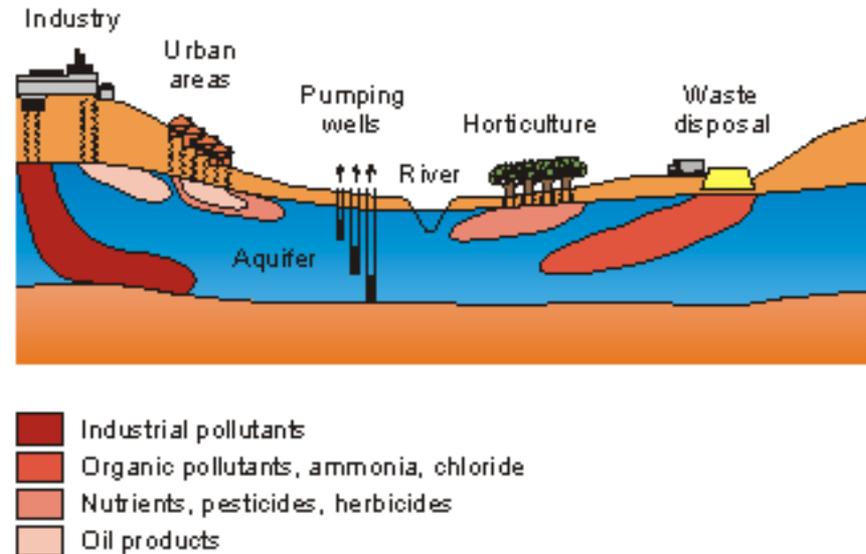
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**Gang Chen, Amy Chan Hilton, and Kamal Tawfiq**

**Department of Civil and Environmental Engineering  
FAMU-FSU College of Engineering**



## Shallow Vadose Zone



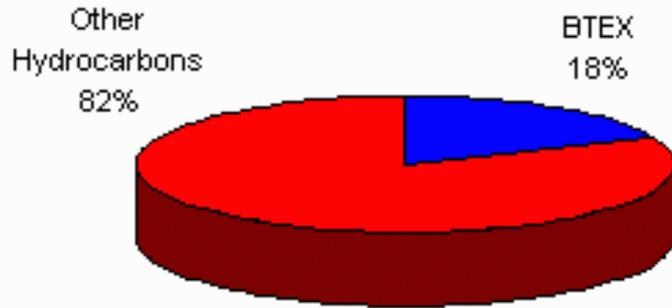
## BTEX, PCE, TCE and Iron Contamination

# Groundwater in Florida

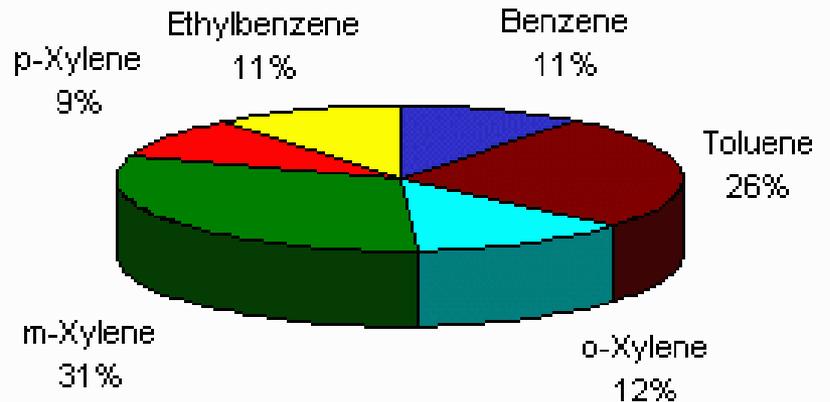


## BTEX Contamination

### Percent BTEX In Gasoline (% weight)



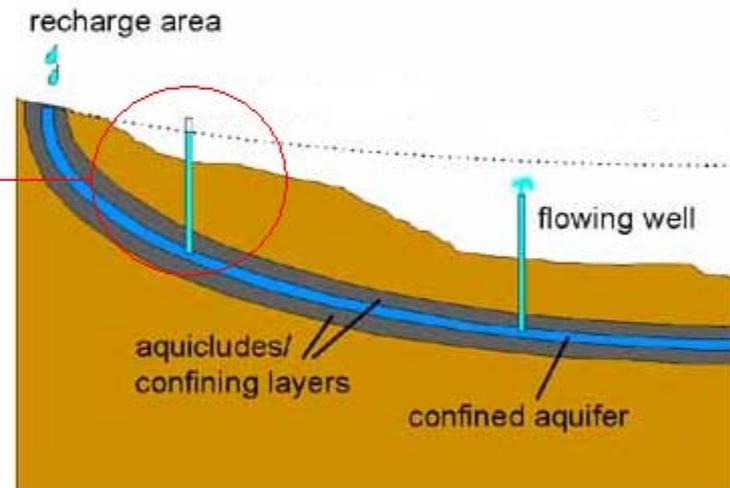
### BTEX Components of Gasoline (% weight)



# Gasoline Composition



- Organic compound degradation
- Volatile organic compound vaporization
- Metal oxidation and precipitation



## Water-Filled Trench

# Objectives

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- **Explore the possibility of the usage of water-filled trench in treating contaminated groundwater**
- **Identify dominating mechanisms**
  - Organic decomposition
  - Volatile organic compound vaporization
  - Metal oxidation and precipitation
- **Investigate the effect**
  - Effect of dissolved oxygen
  - Effect of alkalinity
- **Promote field applications**
  - Quantify removal rate

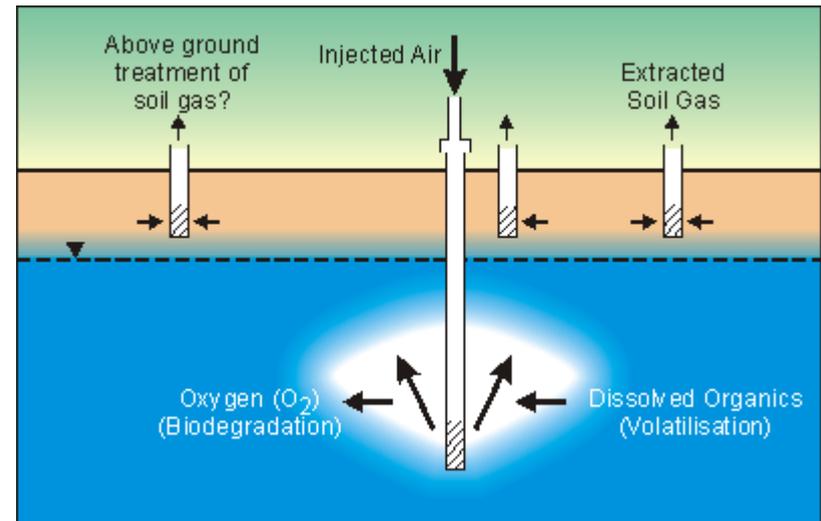
# Outline

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- **Theoretical consideration**
  - Organic compound removal
  - Volatile organic compound removal
  - Metal removal
  - Effect of alkalinity
- **Simulated Groundwater**
- **Water-Filled Trench Experiments**
- **Water-Filled Trench with an Aerobic Filter**
- **Expected Results and Modeling**

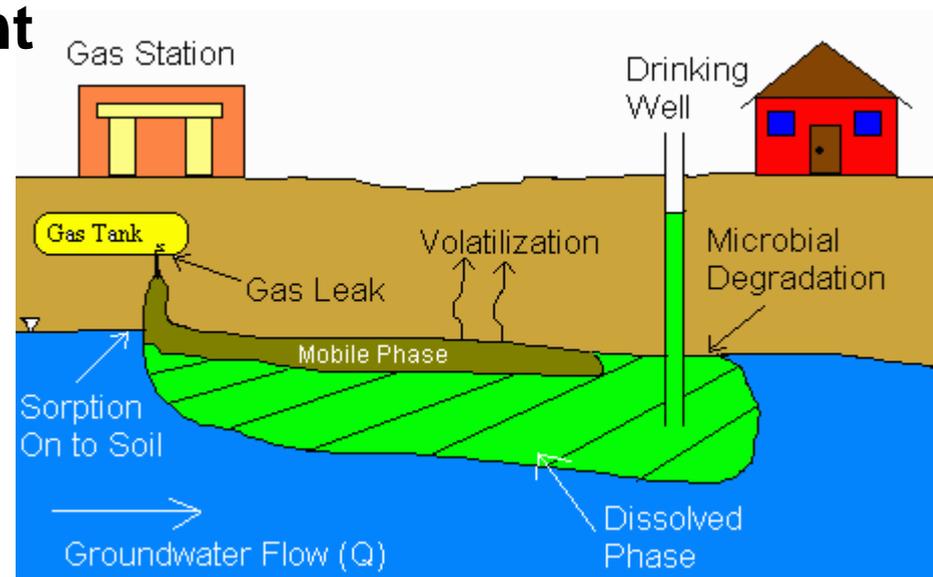
# Organic Decomposition

- **BTEX, pesticides, and polycyclic aromatic hydrocarbons (PAHs)**
- **Aerobic process**
- **Nutrient requirement**
- **pH and temperature**



# Volatile organic compound vaporization

- Volatile organic contaminants
- High vapor pressure and low water solubility
- Temperature-dependent
- BTEX
- PCE and TCE



# Metal Removal

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- Heavy metals to be transformed and precipitate
- pH and redox conditions
- Alkalinity



# Dissolved Oxygen

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- **Organic compound degradation**
- **Electron acceptor**
- **1.0 mg/L DO required for 0.32 mg/L BTEX**
- **Mass transfer between air and groundwater**
- **Heavy metal oxidation and precipitation**

# **Alkalinity**

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- $\text{HCO}_3^-$ ,  $\text{CO}_3^{2-}$  and  $\text{OH}^-$
- **Alkalinity loss during organic degradation,  $\text{CO}_2$  stripping and metal oxidation and precipitation**
- **Most important for heavy metal oxidation and precipitation**
- **Addition of external alkaline agents such as lime**

# Simulated Groundwater

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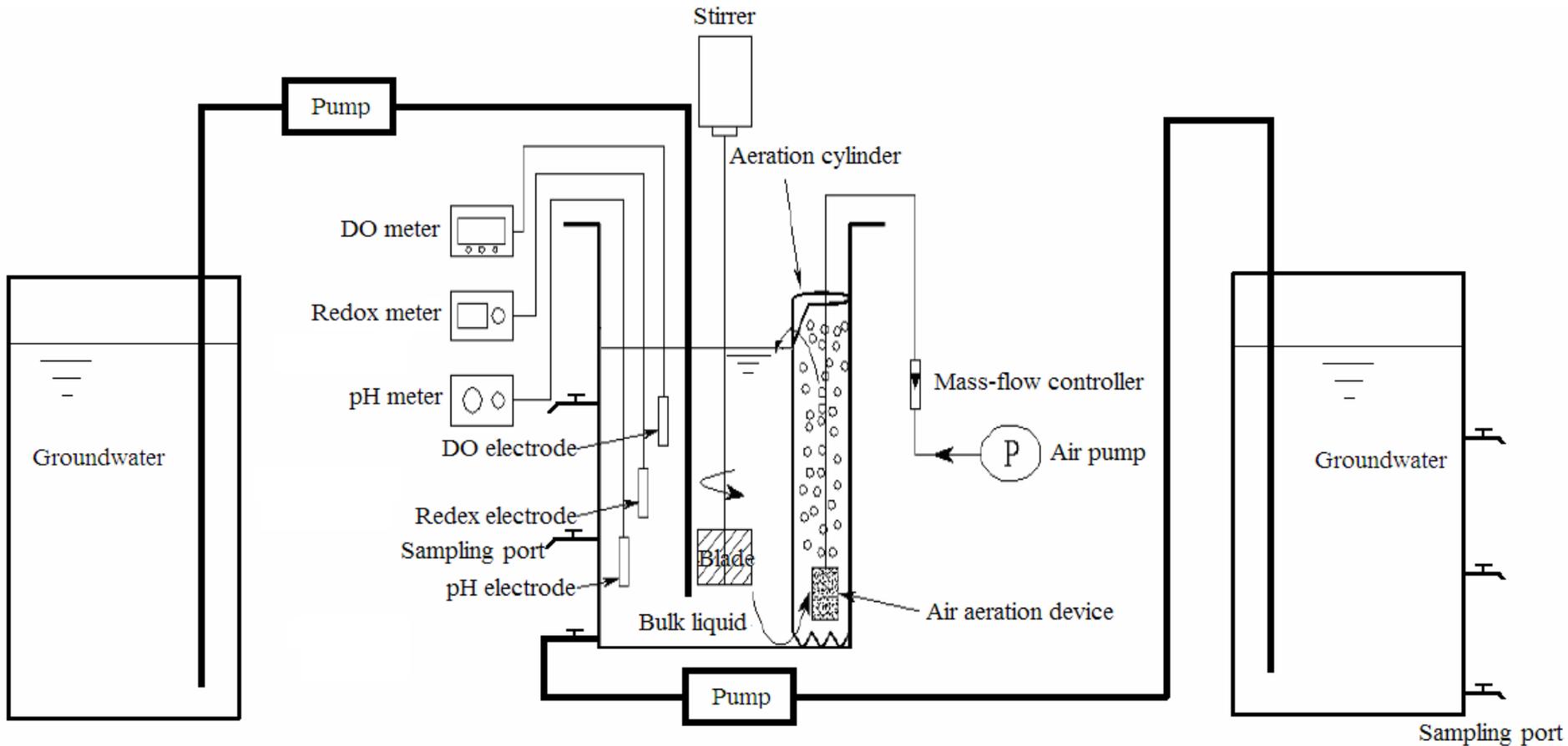
- Benzene, 25 mg/L
- Toluene, 25 mg/L
- Ethylbenzene, 25 mg/L
- *p*-xylene, 25 mg/L
- PCE, 20 µg/L
- TCE, 200 µg/L
- Glucose, 50 mg/L
- Ferrous iron, 10 mg/L



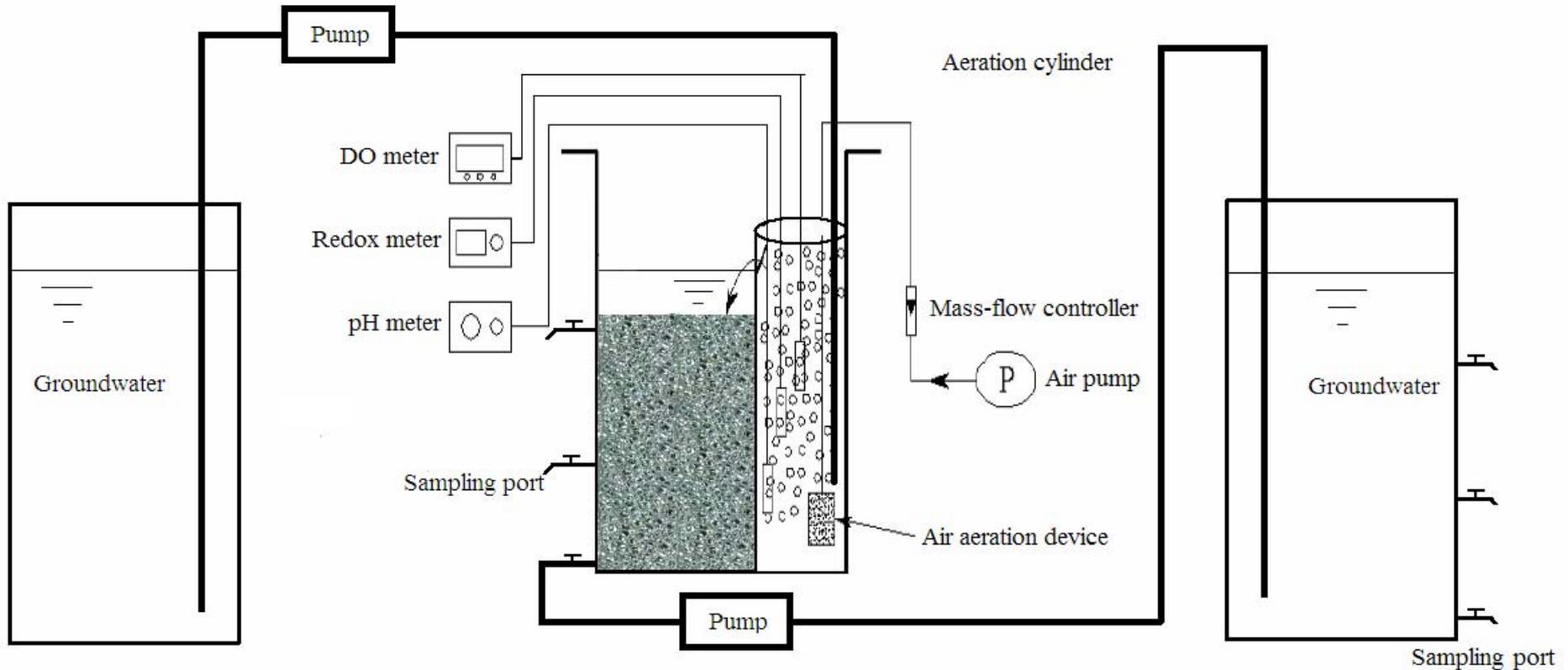
Benzene, toluene, *p*-xylene, PCE, and TCE to be monitored by a GC with a flame ionization detector and helium as the carrier gas

Glucose and iron to be quantified by HPLC and spectrophotometer

# Water-Filled Trench Setup



# Water-Filled Trench with a Filter



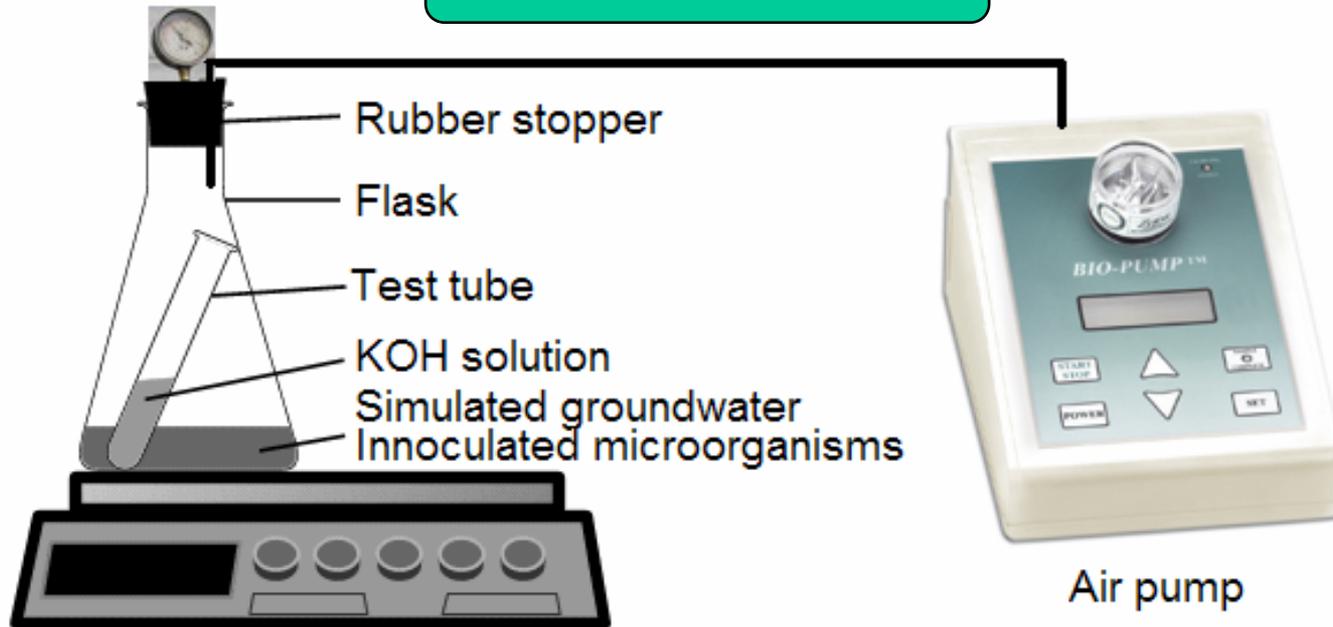
# Other Parameters to Be Monitored

- Redox Potential
- pH
- Conductivity
- Temperature



# Control Experiments

Biological Process?



$$C_p = C^* P \left[ \frac{(1 - P_w/p)(1 - \theta P)}{(1 - P_w)(1 - \theta)} \right]$$

$C_p$ : oxygen concentration at nonstandard pressure

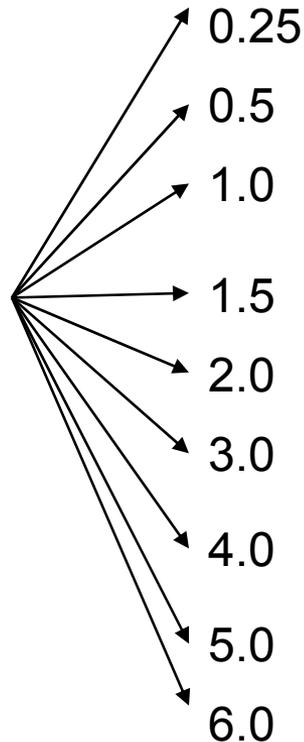
$C^*$ : oxygen concentration at standard pressure

$P$ : nonstandard pressure

$P_w$ : partial pressure of water vapor

# Expected Results — Water-Filled Trench

Benzene, 25 mg/l  
Toluene, 25 mg/l  
xylene, 25 mg/l  
PCE, 20 µg/l  
TCE, 200 µg/l  
Glucose, 50 mg/l  
Fe<sup>2+</sup>, 10 mg/l



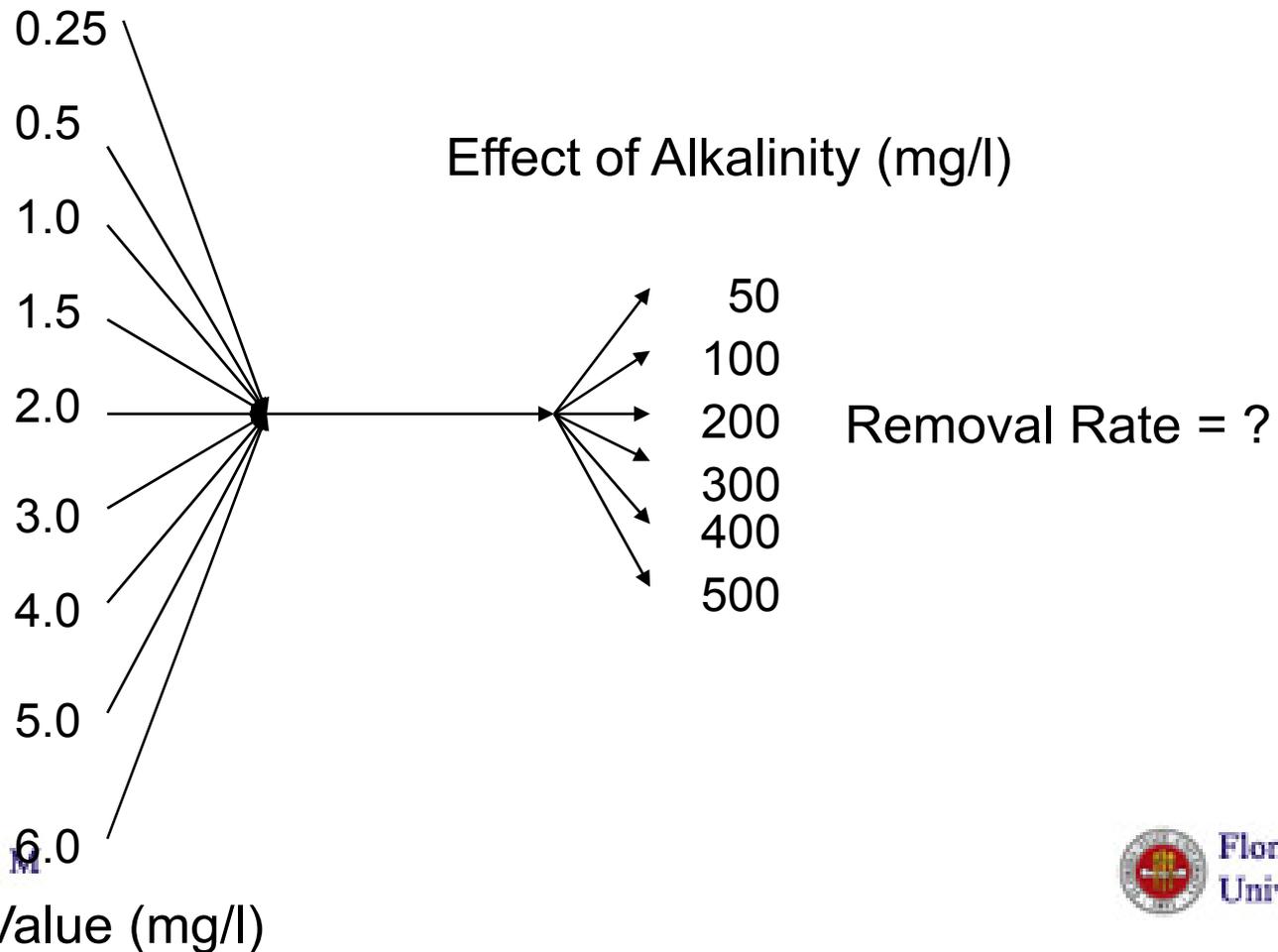
Benzene=? Toluene=?  
Xylene=? Glucose=?  
PCE=? TCE=? Fe<sup>2+</sup>=?

DO Value (mg/l)

pH=? Redox=? Microbial activity=?

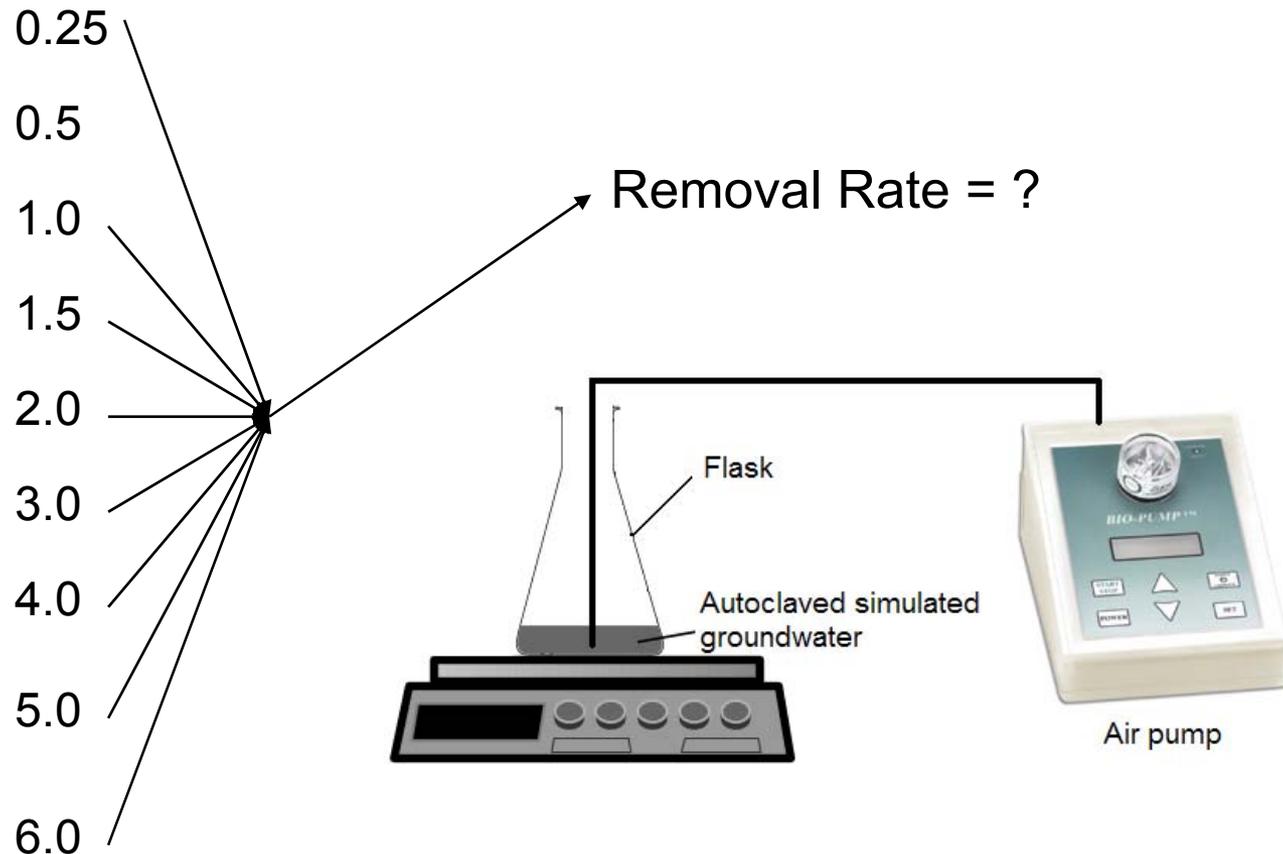
# Expected Results — Water-Filled Trench

Organic Removal — Biological Process?



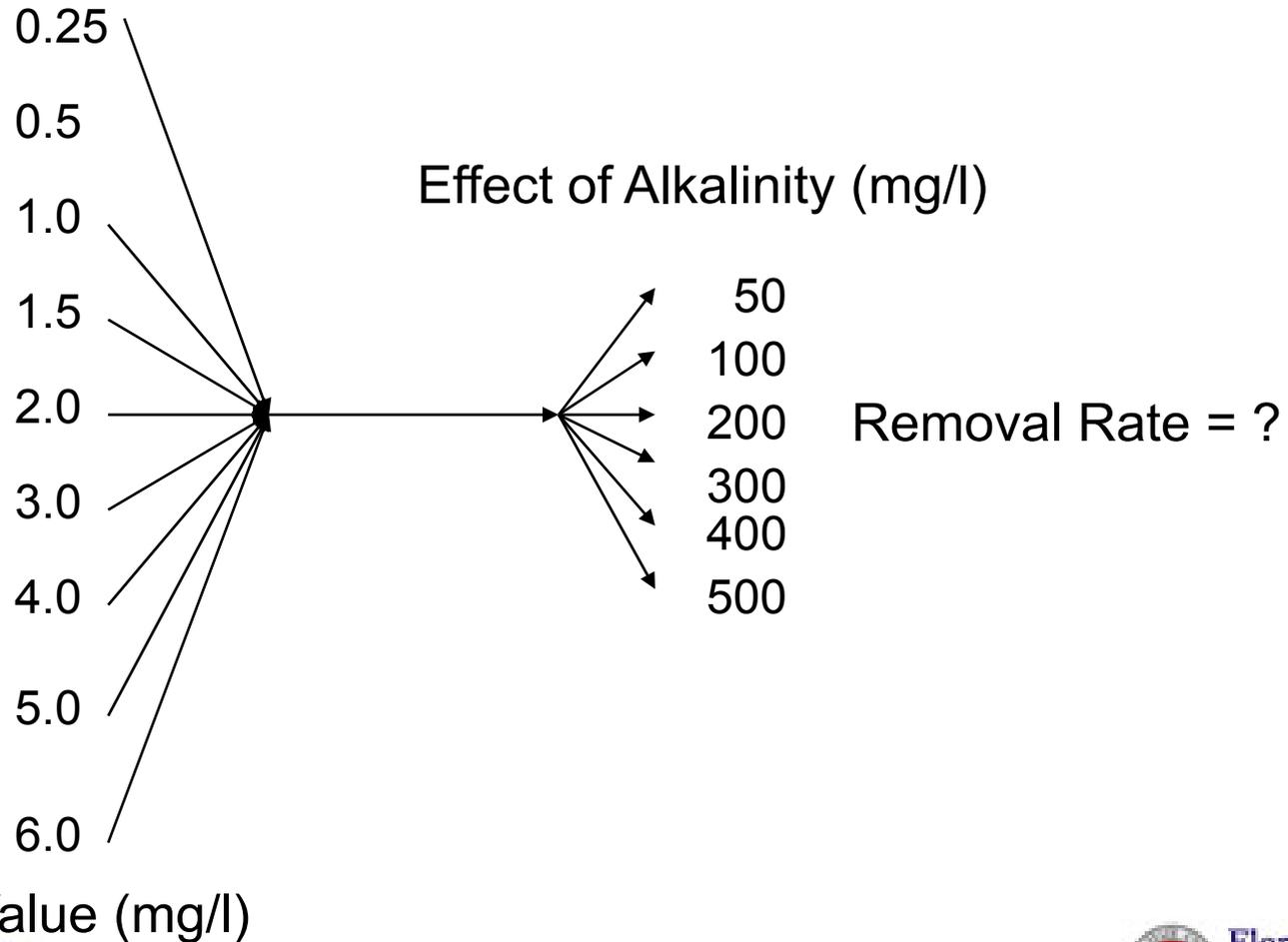
# Expected Results — Water-Filled Trench

BTEX, PCE & TCE Removal — Evaporation?

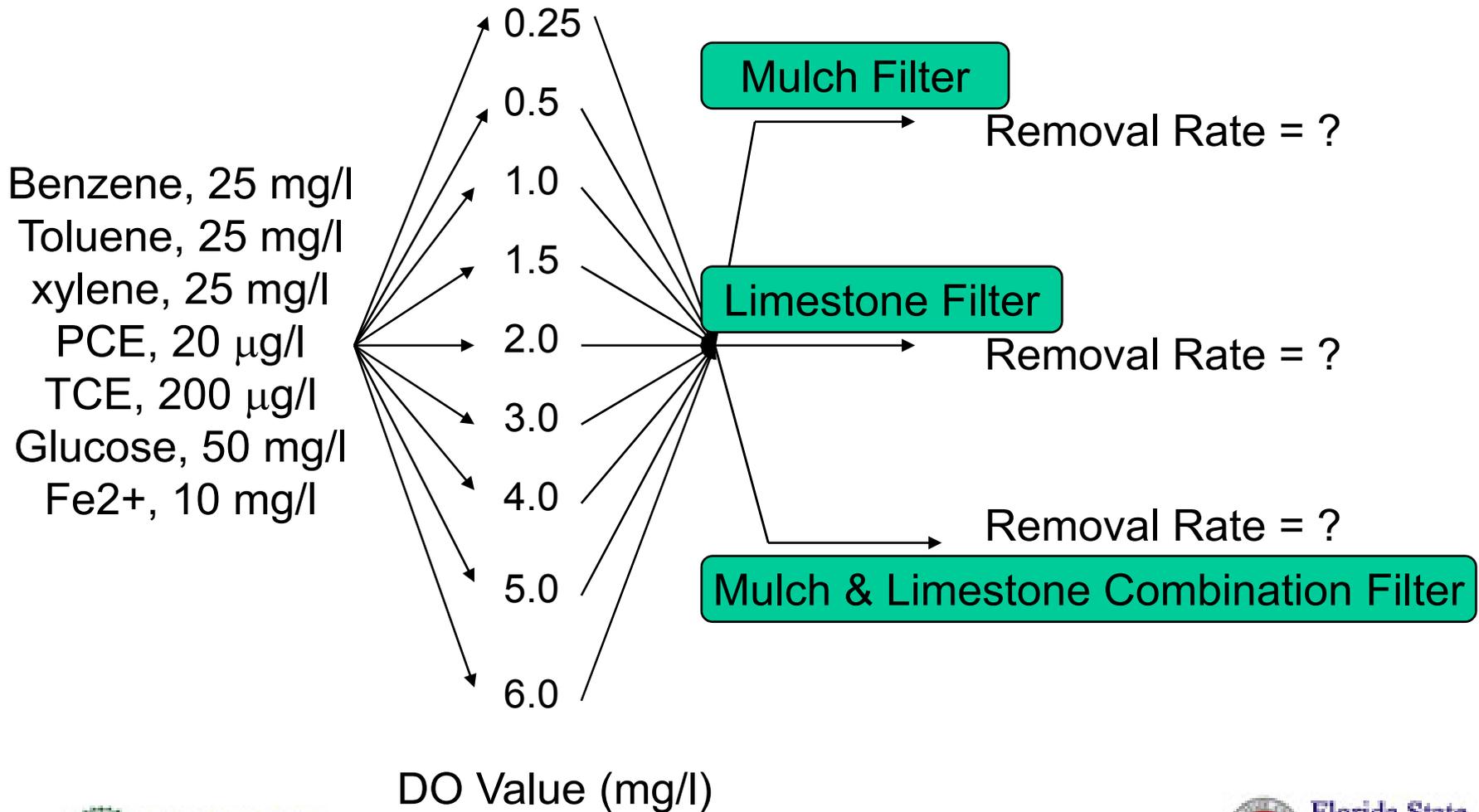


# Expected Results — Water-Filled Trench

Fe<sup>2+</sup> Removal — Biological or Chemical Oxidation and Precipitation?



# Expected Results — Water-Filled Trench with Aerobic Filter



# Biodegradation Simulation

Organic contaminant degradation:  $\frac{dS}{dt} = -\frac{\mu_{\max} S}{K_m + S}$

$$S = K_m W\left[\frac{S_0}{K_m} \exp\left(\frac{S_0 - \mu_{\max} t}{K_m}\right)\right] \quad W(x) + \ln[W(x)] = \ln(x)$$

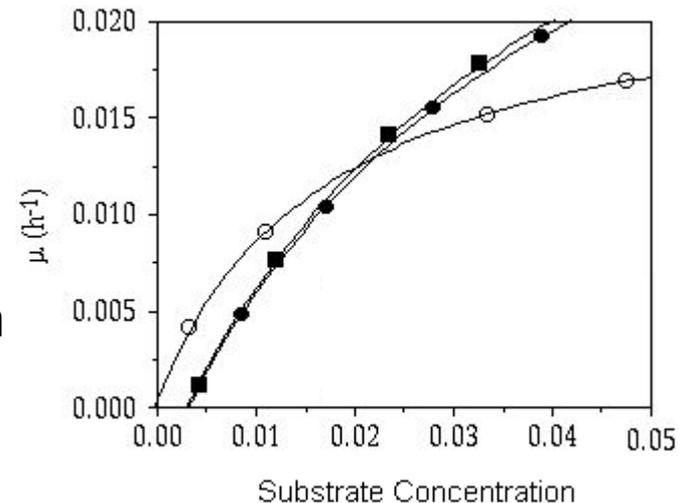
S: organic compound concentration

$\mu_{\max}$ : maximum bacterial growth rate

$K_m$ : half saturation constant

$S_0$ : initial organic compound concentration

W: Lambert function x: argument of W



# Evaporation Simulation

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## Volatile organic contaminants mass transfer:

$$M = K_L a (C_L^* - C_0)$$

M: mass of volatile organic contaminants transferred per unit time and volume

$K_L$ : mass transfer coefficient

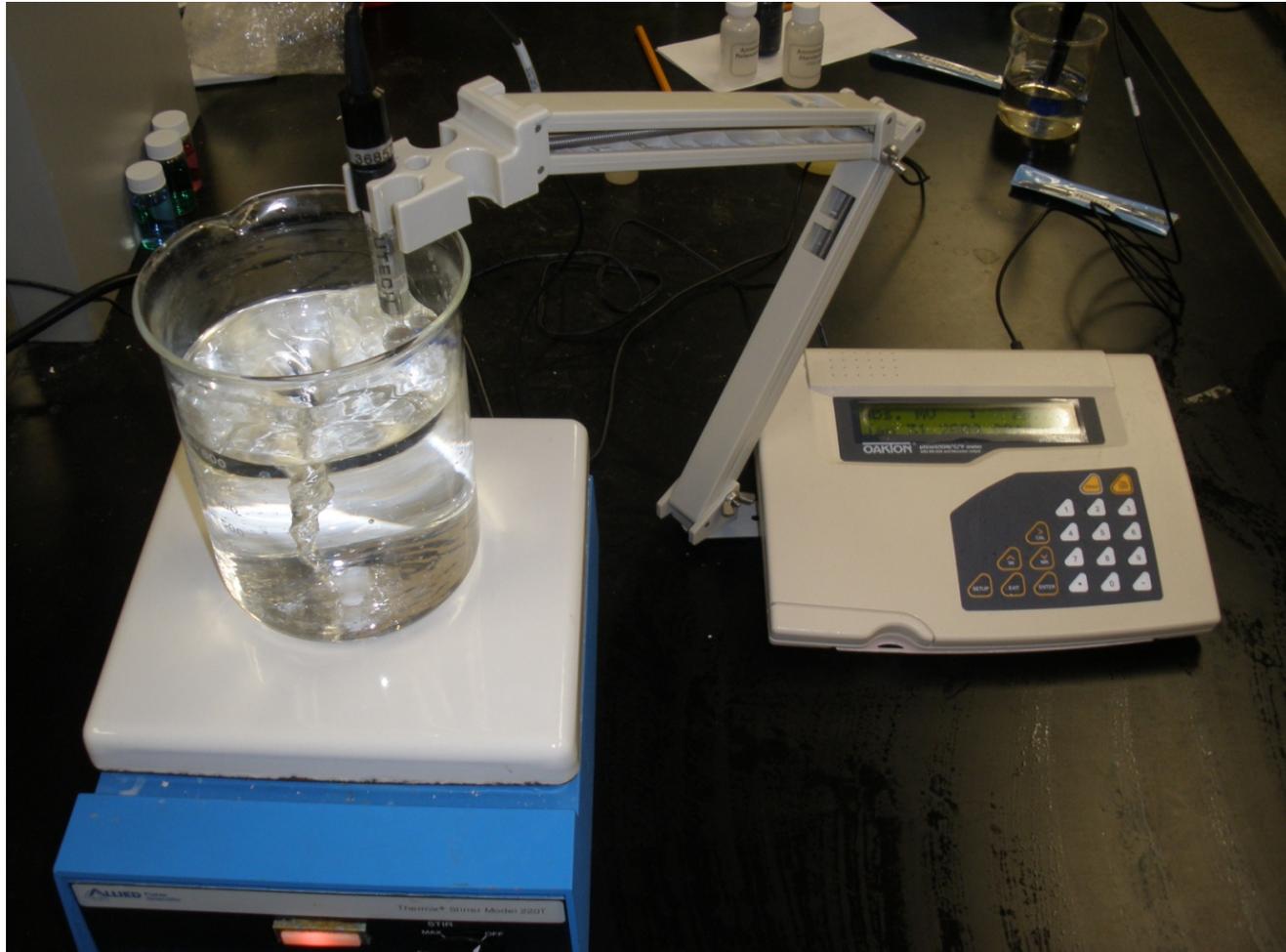
a: effective mass transfer area

$C_L^*$ : liquid phase concentration in equilibrium with gas phase concentration

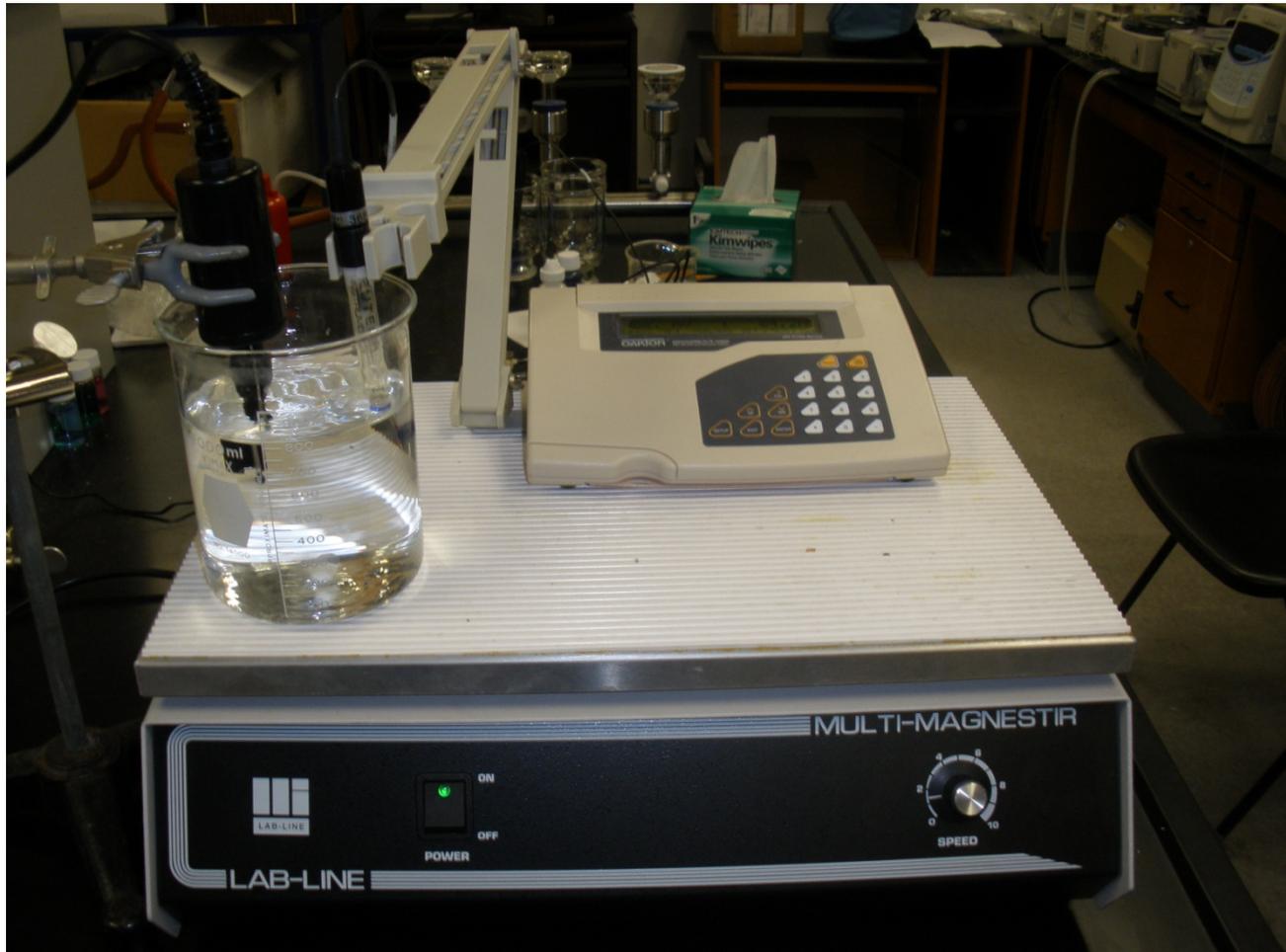
$C_0$ : bulk phase liquid concentration



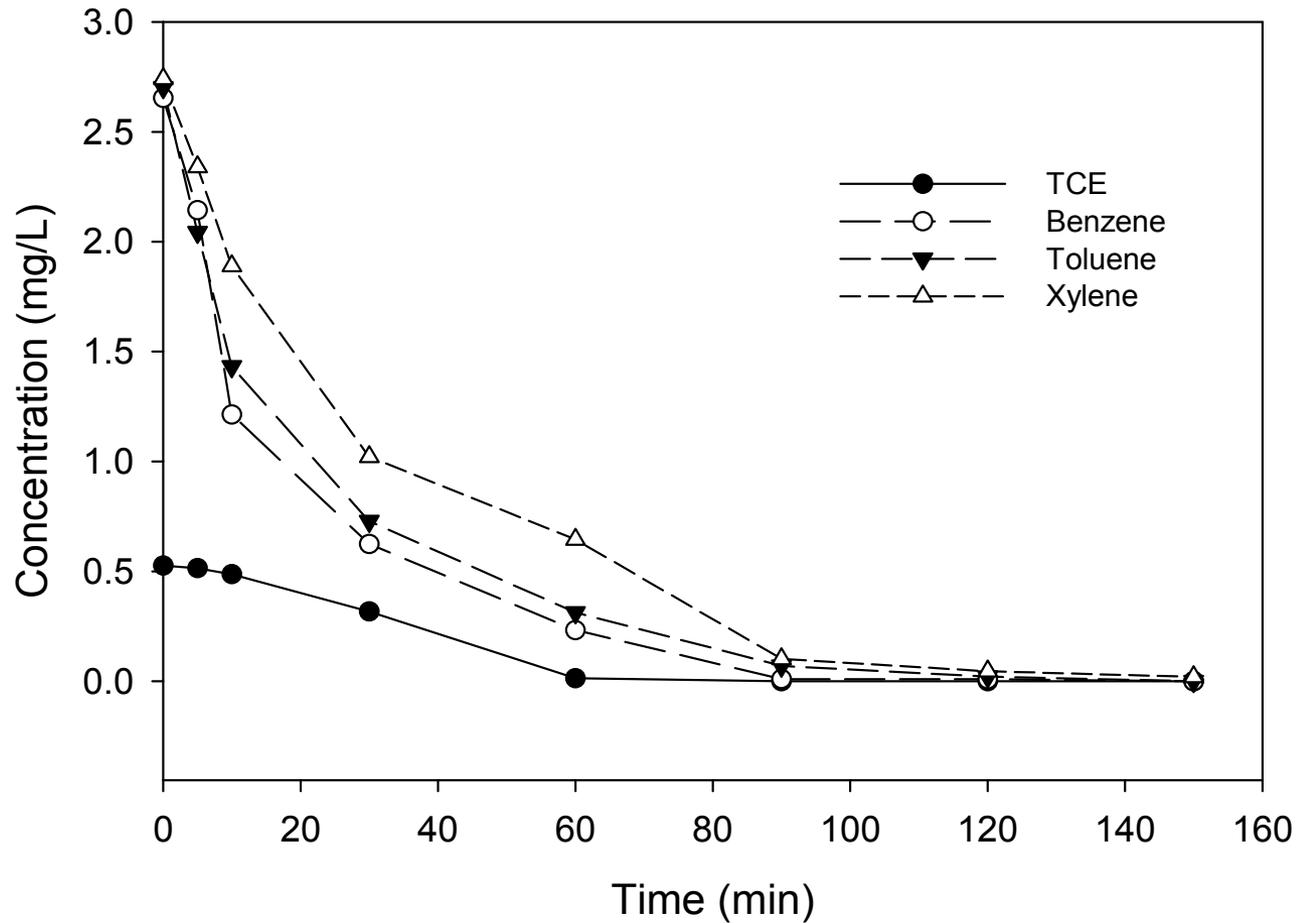
# Evaporation



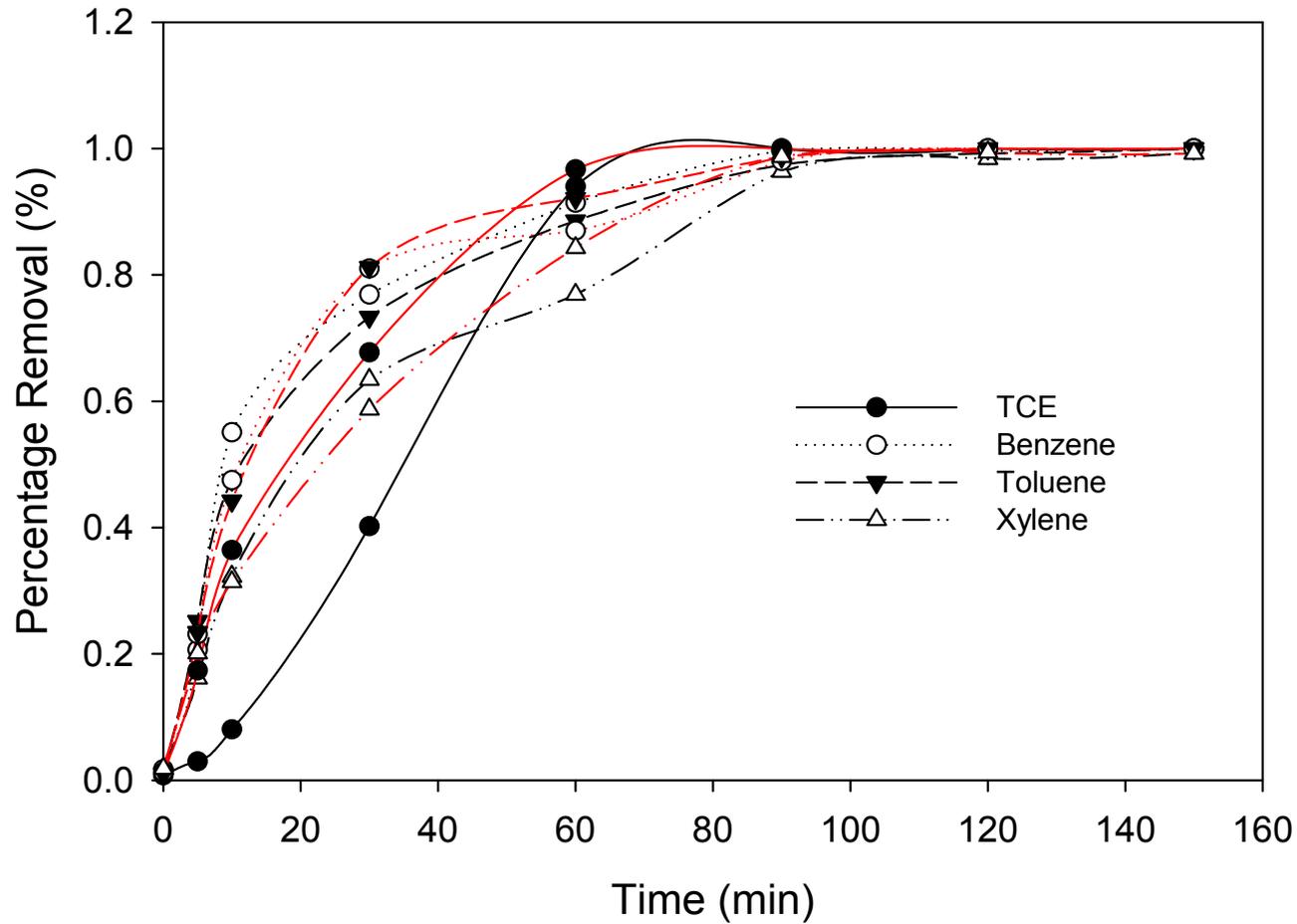
# Evaporation



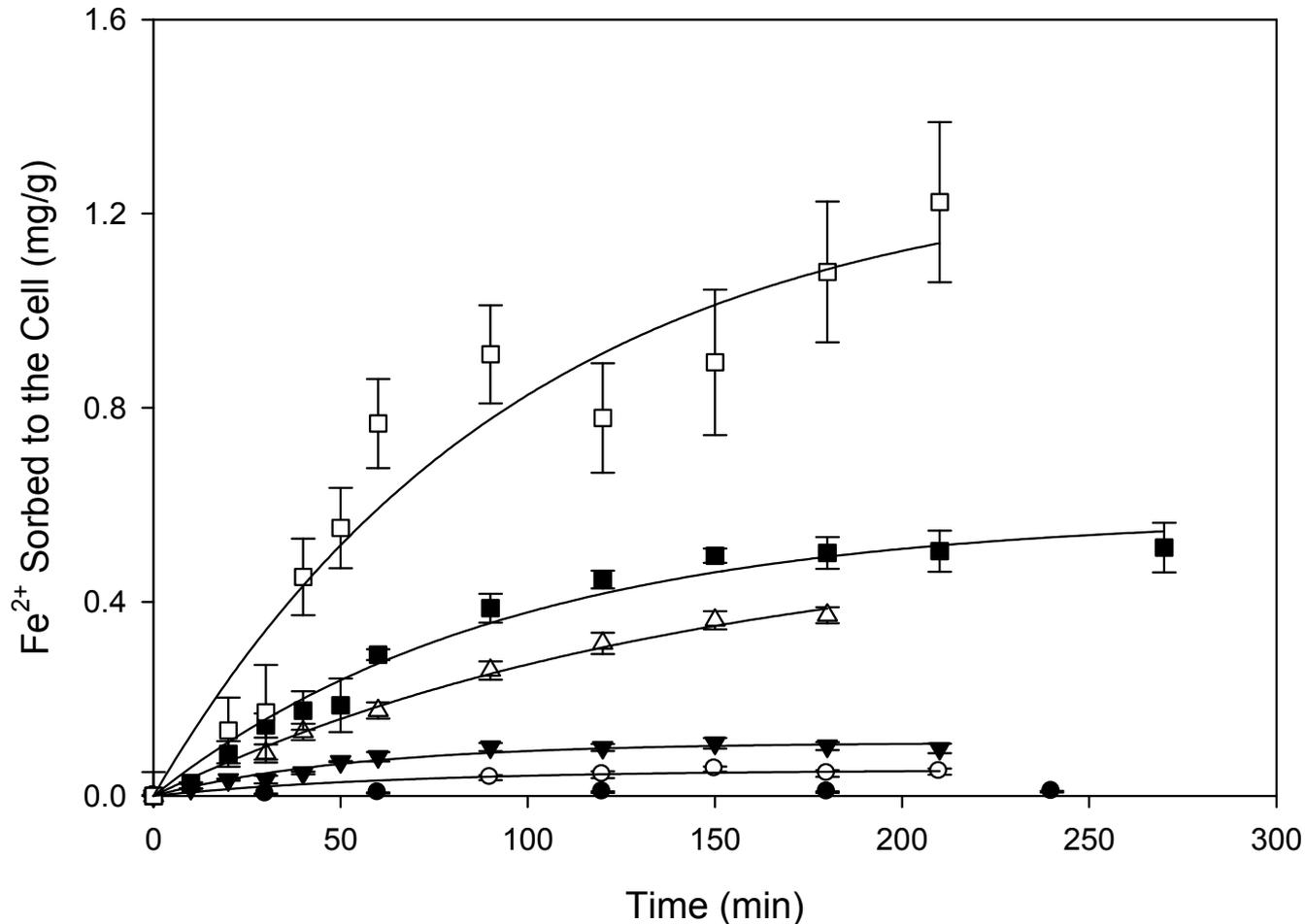
# Evaporation



# Evaporation



# Ferrous Iron Sorption



ferrous iron concentrations  
investigated in this study: 0.25, 2.0,  
3.5, 10.0, 15.0, and 30.0 mg/l

# Ferrous Iron Sorption

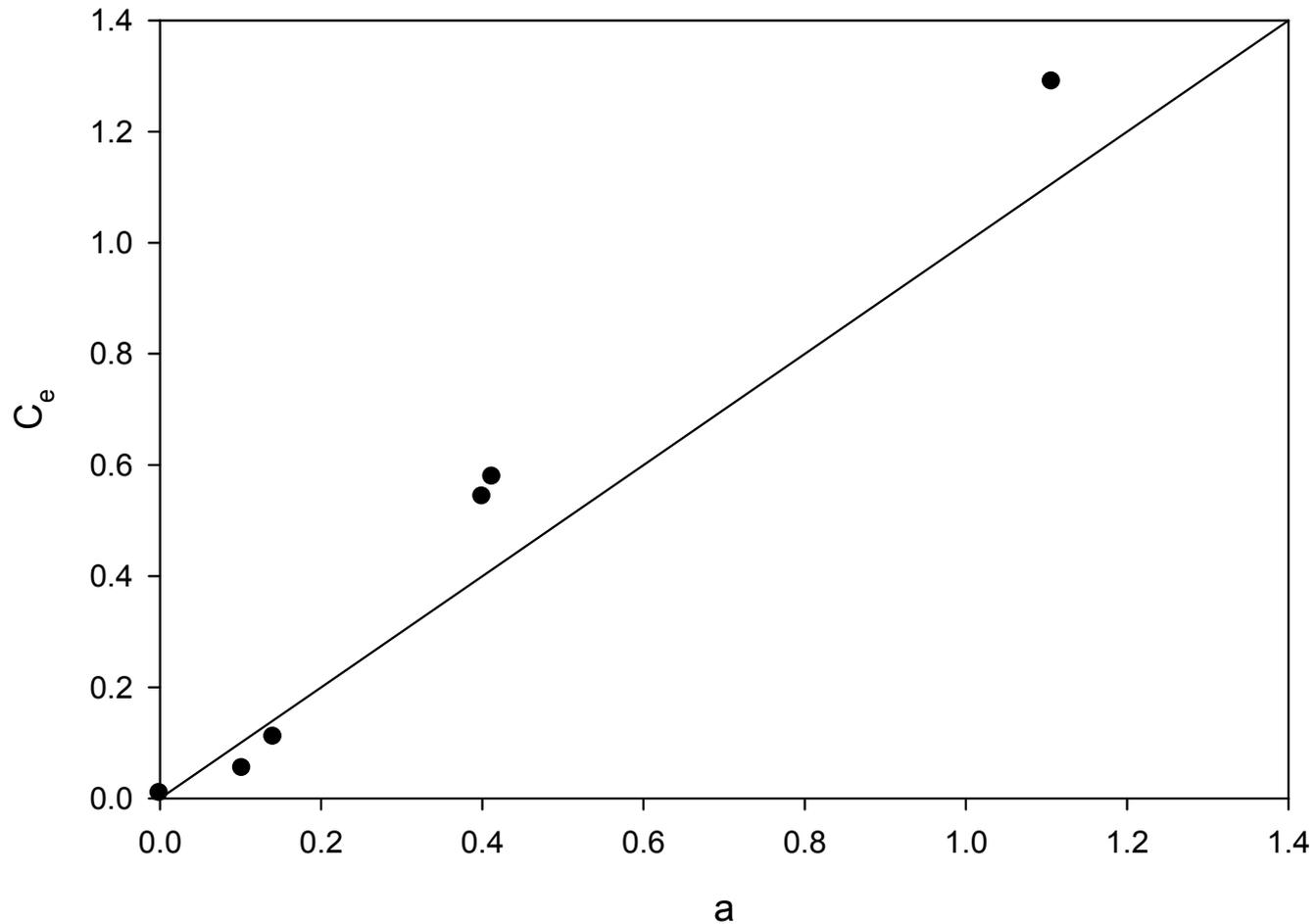
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$$q = a(1 - b^C)$$

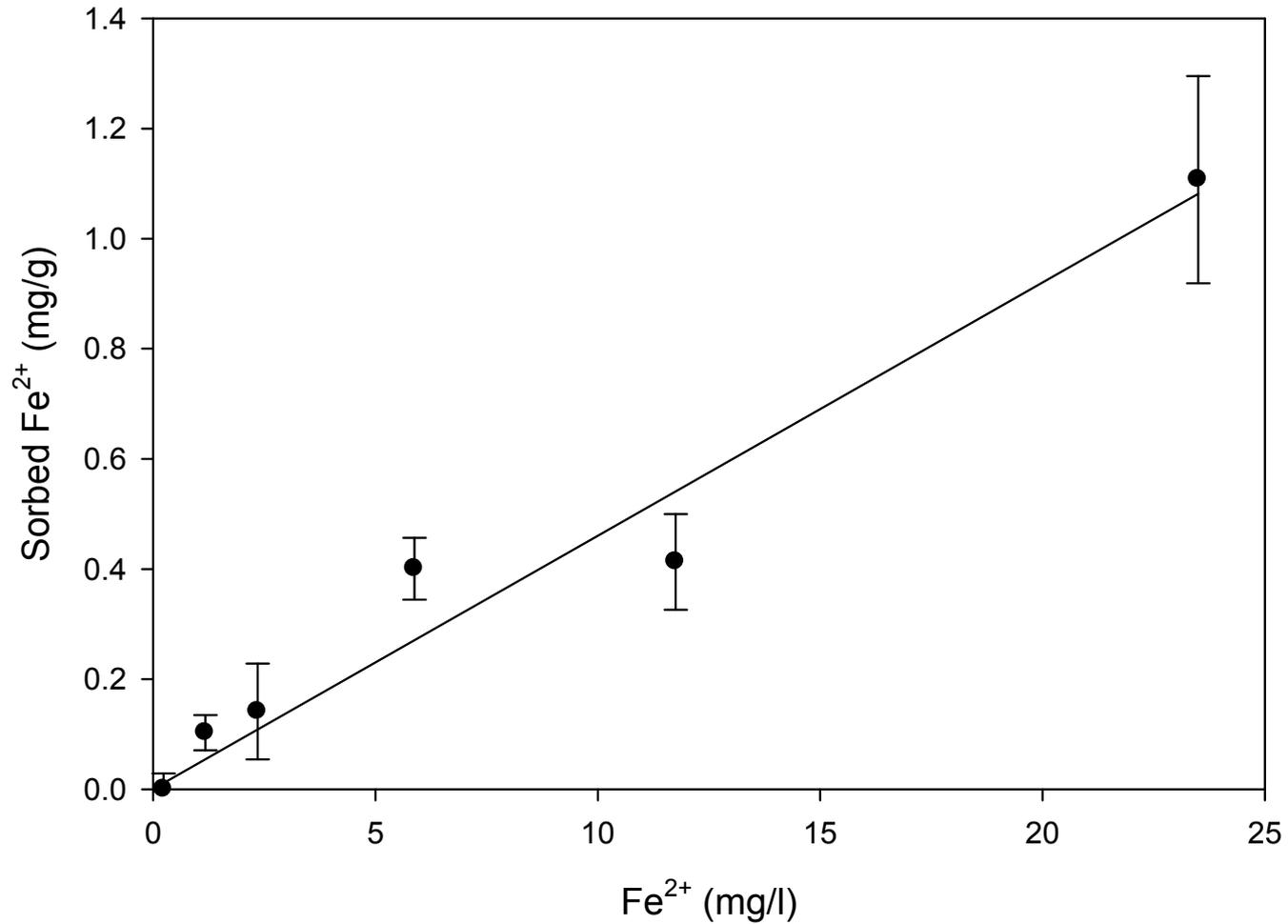
	a	b	R <sup>2</sup>
0.25	0.0085 ± 0.0002	0.9796 ± 0.0013	0.9970
2.00	0.0533 ± 0.0058	0.9850 ± 0.0056	0.9606
3.50	0.1099 ± 0.0058	0.9818 ± 0.0025	0.9609
10.0	0.5424 ± 0.0473	0.9931 ± 0.0010	0.9945
15.0	0.5779 ± 0.0317	0.9894 ± 0.0013	0.9788
30.0	1.2890 ± 0.1837	0.9898 ± 0.0029	0.9214

# Ferrous Iron Sorption

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# Ferrous Iron Sorption



# Questions?

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